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solution. homogeneous mixture made of solvent and solute.
aqueous solution.
solutions with water as the solvent. brownian Motion. chaotic movement of colloidal particles. colloid.

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Terms in this set (47)

characteristics of the

water molecule. 1)

triatomic 2) each O-H

bond is highly polar b/c

of oxygen's high

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electronegativity,
water is a whole is a
polar molecule

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Water & Solutions 6
Chapter 17 Assignment
& Problem Set 24
Honors Explain why
propanol (C₃H₇OH)
will dissolve in both

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gasoline and water 25
Suppose an aqueous
Chapter 17 Additional
Aspects of Aqueous
Equilibria Chapter 17
Additional Aspects of
Aqueous Equilibria
Chemistry, The Central
Science,

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Systems. Milbank High
School; 2 Chapter 17 -
Review. What is the
effect of a wetting
agent on surface
tension? How many
nonbonding pairs of
electrons are in a
water molecule? How
does the surface
tension of water
compare with the
surface tensions of
most ...

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Water and Aqueous

Systems 17.1 - Liquid

Water and Its

Properties • polar

molecule due to bent

shape results in

hydrogen bonds •

hydrogen bonds are

responsible for water's

unusual properties o

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surface tension o high
specific heat capacity
17.2 - Water Vapor and
Ice • relatively high
melting point & boiling
point • expansion of
water as it freezes to
ice results in lower
density 17.3 ...

Chapter 17 Water and Aqueous Systems 171 Liquid Water and ...

Chapter 17 Water And
Aqueous This is
likewise one of the

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AP Chemistry Chapter
17 Additional Aspects
of Aqueous Equilibria -
2 - Sample Exercise
17.1 (p. 720) What is
the pH of a solution
made by adding 0.30
mol of acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) and 0.30 mol
of sodium acetate ($\text{NaC}_2\text{H}_3\text{O}_2$) to enough
water to make 1.0 L of
solution? (4.74)

Practice Exercise 17.1

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Chapter 17. Answers Additional Aspects of Equilibrium

Start studying Chapter 17: Water. Learn vocabulary, terms, and more with flashcards, games, and other study tools. Search. Create. Chapter 17: Water. STUDY. Flashcards. Learn. Write. Spell. ... compounds that do not conduct an electric current in either an

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aqueous solution or the molten state; some very polar molecules are nonelectrolytes until ...

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Aqueous Ionic
Equilibria -- Chapter 17

1. Buffer Solutions A
Buffer Solution is an
acid/base equilibrium
system that is capable
of maintaining a
relatively constant pH
even if a significant

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amount of strong acid or base is added. (a) Components of a buffer solution: a mixture of a weak acid and its conjugate base

Aqueous Ionic Equilibria -- Chapter 17

AP Chemistry Chapter 17 Additional Aspects of Aqueous Equilibria - 1 - Chapter 17.

Additional Aspects of Equilibrium . 17.1 The Common Ion Effect •

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The dissociation of a weak electrolyte is decreased by the addition of a strong electrolyte that has an ion . in common. with the weak electrolyte.

Chapter 17. Additional Aspects of Equilibrium

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For which salt should

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the aqueous solubility be most sensitive to pH? A) MgCl_2 B) $\text{Mg}(\text{NO}_3)_2$... Calculate the pH of a solution prepared by dissolving of acetic acid and of sodium acetate in water sufficient to yield of solution. The K_a ...

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Chapter 17 Notes -
Carbonyl Compounds
Carbonyl Functional
Groups. ...
formaldehyde in water
is mainly
dihydroxymethane ...
(very little gem-diol in
aqueous solution)
reaction rate is slow at

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pH 7 but faster in acid
or in base (catalysis)
Base-Catalyzed
Hydration. OH^- is a
strong nucleophile Acid-
Catalyzed Hydration.

Chapter 17 notes - Portland State University

Chapter 17 (Additional
Aspects of Aqueous
Equilibria) - Part 3 -
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views. 42:02. Chapter
8 - Basic Concepts of

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Chemical Bonding: Part
7 of 8 - Duration: 6:41.

...

Chapter 17 - Additional Aspects of Aqueous Equilibria: Part 6 of 21

Chapter 17: Solubility
& Complex Ion

Equilibria The goal of
this chapter is to
understand the
equilibria that exist
between ionic solids
and their ions in

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solution, and factors that affect that equilibrium. write (heterogeneous) equilibrium equations & K expressions calculate and interpret K_{sp} K_{sp} is the solubility product constant using K

Chapter 17: Overview of the Chapter Solubility & Complex ...

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Aqueous Ionic
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17.1 Extent of Acid-

Base Reactions:

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Simulation In Chapter 5, you learned that acids and bases react to form water and a salt and that these reactions are called neutralization reactions because, on completion of the reaction, the solution is neutral. As shown in Table 17.1, however, acid-base reactions do not

CHAPTER 17: **Advanced Acid-Base**

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